

The History & Make-Up of Atoms

Atoms

Atoms are the basic building blocks for all objects in universe, and all elements discovered (or made) are made of different atoms (by elements, I mean “Carbon”, Helium”, Aluminum”, etc.) The atom was originally thought to be smallest particle around, but then, discoveries of sub-atomic particles were made! We have:

1. Proton (+) = **Defines the element!**, positively charged, mass = 1.7×10^{-27} kilograms
2. Electron (-) = negatively charged, mass = 9.1×10^{-31} kilograms
3. Neutron = no charge, same mass as proton

Math Problem!

The electron’s mass is considered negligible relative to the proton. Why?
Show with examples or a proof.

This is a COMPARATIVE size question, so we just use the exponents as a ratio to compare!
 10^{-27} vs $10^{-31} \rightarrow 10^{-27} / 10^{-31} \rightarrow 10 (-27-(-31)) = 10^{(-27+31)} = 10^4 = 10000$. The proton is ten thousand times more massive than the electron!

Atom Make-up

Protons & Neutron = exist in nucleus

Electron = exist outside of the nucleus (more on this later)

The Periodic Table

Elements are organized into the Periodic Table of Elements. They are organized into columns by their similarities in chemical properties:

Periodic Table of the Elements

1 IA 11A																	18 VIII A 8A
1 H Hydrogen 1.008																	2 He Helium 4.003
3 Li Lithium 6.941	4 Be Beryllium 9.012											5 B Boron 10.811	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180
11 Na Sodium 22.990	12 Mg Magnesium 24.305	3 IIIB 3B	4 IVB 4B	5 VB 5B	6 VIB 6B	7 VIIB 7B	8 VIII 8	9 VIII 8	10 VIII 8	11 IB 1B	12 IIB 2B	13 Al Aluminum 26.982	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.066	17 Cl Chlorine 35.453	18 Ar Argon 39.948
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.88	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.933	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.39	31 Ga Gallium 69.732	32 Ge Germanium 72.61	33 As Arsenic 74.922	34 Se Selenium 78.09	35 Br Bromine 79.904	36 Kr Krypton 84.80
37 Rb Rubidium 84.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.94	43 Tc Technetium 98.907	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.71	51 Sb Antimony 121.760	52 Te Tellurium 127.6	53 I Iodine 126.904	54 Xe Xenon 131.29
55 Cs Cesium 132.905	56 Ba Barium 137.327	57-71 Lanthanide Series	72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.85	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.22	78 Pt Platinum 195.08	79 Au Gold 196.967	80 Hg Mercury 200.59	81 Tl Thallium 204.383	82 Pb Lead 207.2	83 Bi Bismuth 208.980	84 Po Polonium [208.982]	85 At Astatine 209.987	86 Rn Radon 222.018
87 Fr Francium 223.020	88 Ra Radium 226.025	89-103 Actinide Series	104 Rf Rutherfordium [261]	105 Db Dubnium [262]	106 Sg Seaborgium [266]	107 Bh Bohrium [264]	108 Hs Hassium [269]	109 Mt Meitnerium [268]	110 Ds Darmstadtium [269]	111 Rg Roentgenium [272]	112 Cn Copernicium [277]	113 Uut Ununtrium unknown	114 F1 Flerovium [289]	115 Uup Ununpentium unknown	116 Lv Livermorium [298]	117 Uus Ununseptium unknown	118 Uuo Ununoctium unknown
		57 La Lanthanum 138.906	58 Ce Cerium 140.115	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.24	61 Pm Promethium 144.913	62 Sm Samarium 150.36	63 Eu Europium 151.966	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.50	67 Ho Holmium 164.930	68 Er Erbium 167.26	69 Tm Thulium 168.934	70 Yb Ytterbium 173.04	71 Lu Lutetium 174.967	
		89 Ac Actinium 227.028	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium 237.048	94 Pu Plutonium 244.064	95 Am Americium 243.061	96 Cm Curium 247.070	97 Bk Berkelium 247.070	98 Cf Californium 251.080	99 Es Einsteinium [254]	100 Fm Fermium 257.095	101 Md Mendelevium 258.1	102 No Nobelium 259.101	103 Lr Lawrencium [262]	

Alkali Metal

Alkaline Earth

Transition Metal

Semimetal

Nonmetal

Basic Metal

Halogen

Noble Gas

Lanthanide

Actinide

For each element, we can directly relate the amount of protons, electrons and neutrons that exist.
But first, we need to learn some terms!

Symbol of Element

1 or 2 letter abbreviation for each element

Mass Number

- Not always a whole number (more on this later!)
- #protons + #neutrons

Atomic Number

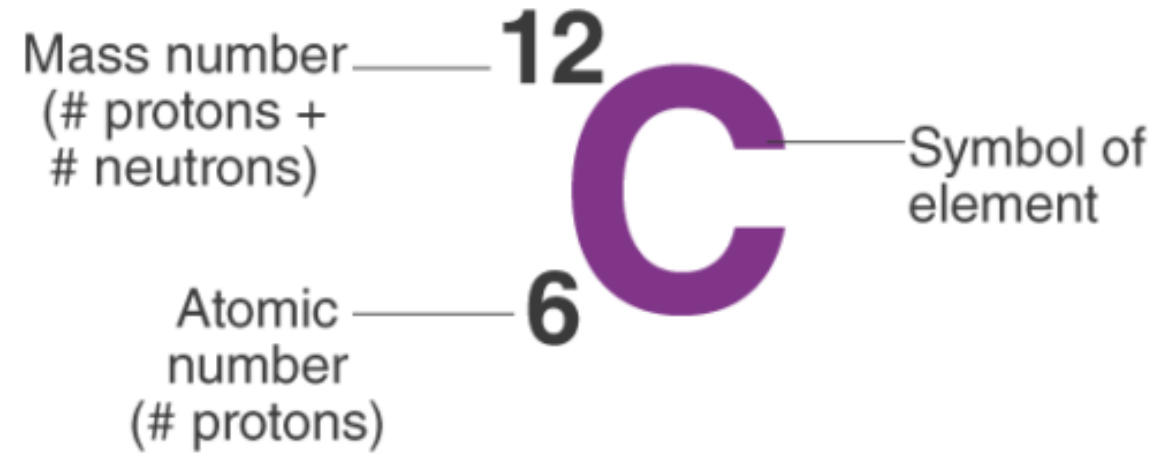
#protons (defines the element!)

Thus,

of protons = atomic number ([defines the element!](#))

of electrons = # of protons (if neutral)

of neutrons = Mass Number - Atomic number



Math Problem!

Using the provided periodic table, calculate the number of protons, electrons, and neutrons for each of the following:

1. Fluorine (F)
2. Iron (Fe)
3. Charged Oxygen (O^{-2})
4. Chlorine (Cl)

1		2												3	4	5	6	7	0			
				Key																1 H hydrogen 1	4 He helium 2	
relative atomic mass		atomic symbol		name										atomic (proton) number								
7 Li lithium 3	9 Be beryllium 4											11 B boron 5	12 C carbon 6	14 N nitrogen 7	16 O oxygen 8	19 F fluorine 9	20 Ne neon 10					
23 Na sodium 11	24 Mg magnesium 12											27 Al aluminium 13	28 Si silicon 14	31 P phosphorus 15	32 S sulfur 16	35.5 Cl chlorine 17	40 Ar argon 18					
39 K potassium 19	40 Ca calcium 20	45 Sc scandium 21	48 Ti titanium 22	51 V vanadium 23	52 Cr chromium 24	55 Mn manganese 25	56 Fe iron 26	59 Co cobalt 27	59 Ni nickel 28	63.5 Cu copper 29	65 Zn zinc 30	70 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	80 Br bromine 35	84 Kr krypton 36					
85 Rb rubidium 37	88 Sr strontium 38	89 Y yttrium 39	91 Zr zirconium 40	93 Nb niobium 41	96 Mo molybdenum 42	[98] Tc technetium 43	101 Ru ruthenium 44	103 Rh rhodium 45	106 Pd palladium 46	108 Ag silver 47	112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	128 Te tellurium 52	127 I iodine 53	131 Xe xenon 54					
133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57	178 Hf hafnium 72	181 Ta tantalum 73	184 W tungsten 74	186 Re rhenium 75	190 Os osmium 76	192 Ir iridium 77	195 Pt platinum 78	197 Au gold 79	201 Hg mercury 80	204 Tl thallium 81	207 Pb lead 82	209 Bi bismuth 83	[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86					
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[277] Hs hassium 108	[268] Mt meitnerium 109	[271] Ds darmstadtium 110	[272] Rg roentgenium 111	Elements with atomic numbers 112 – 116 have been reported but not fully authenticated											

* The Lanthanides (atomic numbers 58 – 71) and the Actinides (atomic numbers 90 – 103) have been omitted.

Relative atomic masses for **Cu** and **Cl** have not been rounded to the nearest whole number.

Solutions

Fluorine (F)

Protons = atm # = 9

Neutrons = $19 - 9 = 10$

Electrons = protons = 9

Iron (Fe)

Protons = atm # = 26

Neutrons = $56 - 26 = 30$

Electrons = protons = 26

Charged Oxygen (O⁻²)

Protons = atm # = 8

Neutrons = $16 - 8 = 8$

2 more electrons = 10 electrons

Chlorine (Cl)

Protons = atm # = 17

Neutrons = $35.5 - 17 = 18.5$

Electrons = protons = 17

Is there more than one answer possible for #3? Why or Why not?

Mathematically yes, BUT if the proton number changes, then we no longer have Oxygen, so there is only the one answer possible (above).

What do you notice about Chlorine?

½ neutron! ½ neutron! Is that possible? No, so see next page ;)

The Truth About The Atomic Mass Number!

Many elements occur naturally in different varieties. As we saw with problem 3, **electrons** may be added or taken away to create charged elements called **ions** (positively charged = cations; negatively charged = anions).

But we can also vary the number of neutrons in the nucleus while **NOT** changing the number of protons (**why is this?**). This creates the same element with different masses and thus different atomic mass numbers. These are referred to as **isotopes** of an element.

Isotopes = Different version of the same element due to its neutrons. They are found in nature in specified %'s (done so experimentally).

For example:

C-12 = Carbon 12 features 6 protons + 6 neutrons in its nucleus; It's Percentage Abundance is 98.90%

C-13 = Carbon 13 features 6 protons + 7 neutrons in its nucleus; It's Percentage Abundance is 1.10%

Carbon's listed and PT table mass is **12.011**. How did that number get calculated?

Via **Weighted Average** Calculations!

(Mass of X isotope x % abundance) + (Mass of Y isotope x % abundance) + = **avg mass** (also referred to as amu)

Math Problems!

1) Set-up the equation to calculate the average atomic mass of Nitrogen (N) based on the information given:

Isotope	Mass	% Abundance
N-14	14.003074	99.63%
N-15	15.000108	0.37%

2) The final grade for “Math Taught the Right Way (MTRW)” is calculated via weighted averages. What is final grade if the following were true?

Homework	Attendance	Final
20% of grade	20% of grade	60% of grade
800 points out of 1000 total points available	16 classes attended out of 20 classes given	90% on test

Answers!

1)

$$(14.003074)(.9963) + (15.000108)(.0037) = 14.007$$

2)

$$.20 (800/1000) + .20 (16/20) + .6 (90/100)$$

$$.20 (4/5) + .20 (4/5) + .6 (9/10)$$

$$.20 (.8) + .20 (.8) + .6 (.9)$$

$$.16 + .16 + .54$$

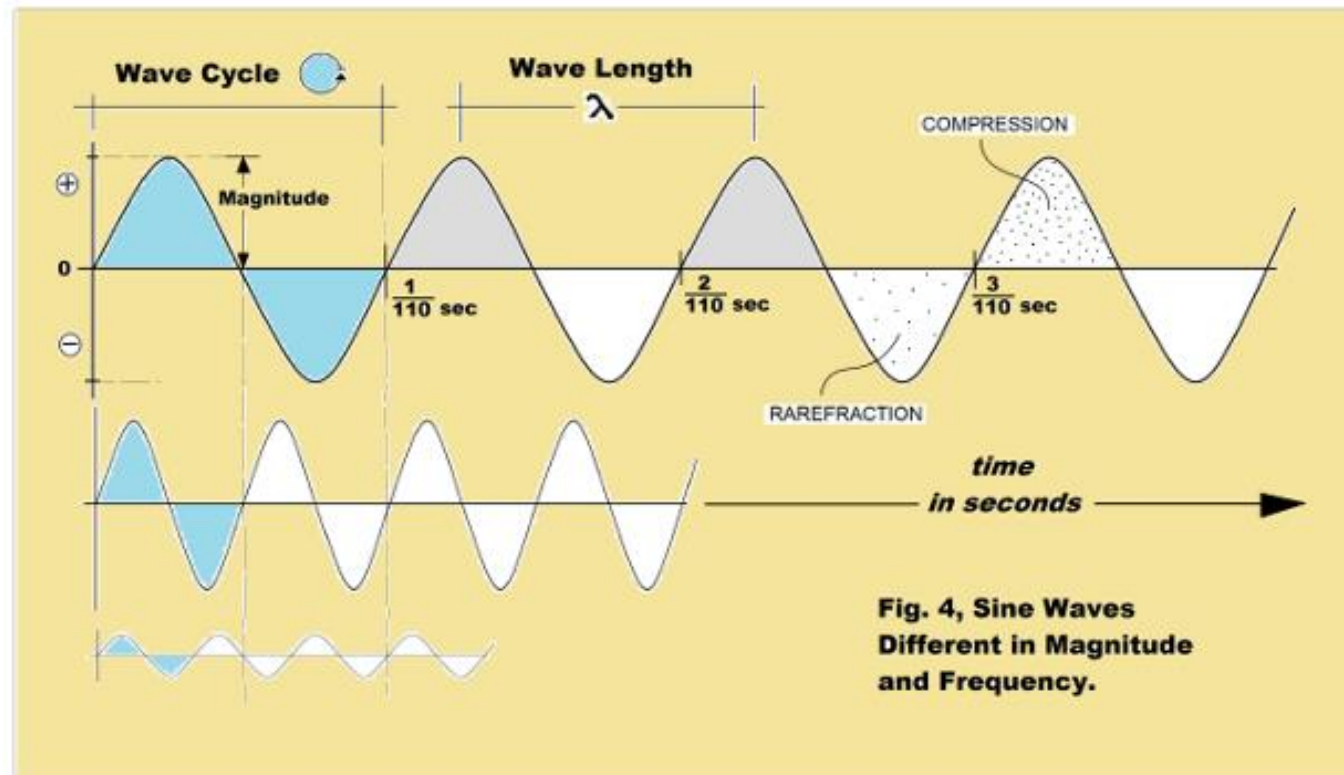
$$.86 = 86\%$$

Let's transition to light and PHOTONS! We will relate this soon to Electrons ;)

Light is defined as both a WAVE and a PARTICLE!

As a PARTICLE, light exists in defined quantities known as Light Quanta or Photons, and these photons have energy associated with them (same is true for ANY moving object – think physics!). A photon is considered to be massless with no electric charge.

As a WAVE, light has properties of waves, including frequency (how often the cycle occurs) and wavelength (the length of each cycle). Mathematically, it looks like this:



Because light travels and has energy, we can calculate it's energy and properties!

$$E = hf$$

E = Energy of the Photon (Joules)

h = Planck's Constant = 6.626×10^{-34} J-s = 6.6×10^{-34} J-s

f = frequency of photon (hertz, cycles/second = 1/s))

BUT, for waves, we can relate the frequency to the wavelength via it's speed (true for any constant wave)!

Speed = frequency (f) x wavelength (λ)

(wavelength = λ (lambda), measured in meters)

For **light** then:

Speed of light = frequency (f) x wavelength (λ)

c = frequency (f) x wavelength (λ)

$$c = f \times \lambda$$

Where c = speed of light = 3×10^8 meters/second – Super fast!

Therefore, combining this equation with the equation above:

$$E = hf = hc/\lambda$$

Why does this matter?

1. Properties of different light types can be studied!
2. Fun with exponent math!

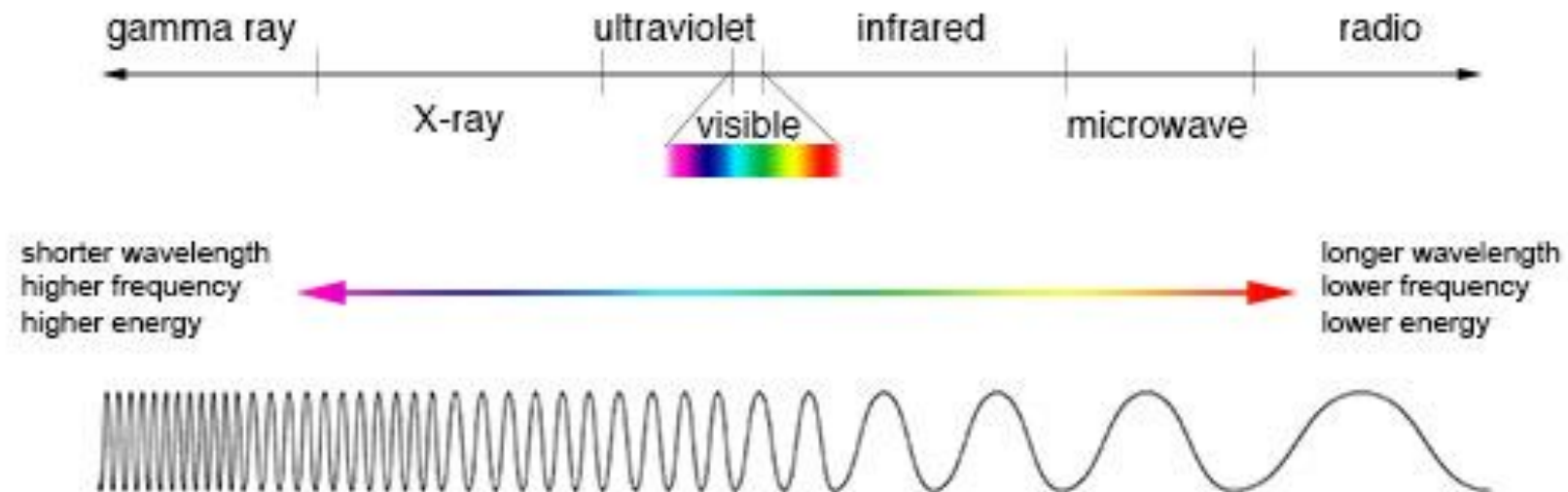
$$E = hf = hc/\lambda$$

Energy and frequency are DIRECTLY related

Energy and wavelength are INVERSELY related

↑ Energy = ↑ frequency = ↓ wavelength

↓ Energy = ↓ frequency = ↑ wavelength



All light types have specified ranges for frequency and wavelength. Commonly, wavelength is used to describe light. The light we see, **visible light**, has wavelengths of 4×10^{-7} meters to 7×10^{-7} meters.

To measure visible light, we normally use nanometers:

1 meter = 1×10^9 nanometers

